

Catalyst

Changes the rate of a chemical reaction without being used up in the reaction

Heterogenous catalysis

Reactant(s) and catalyst are in different phases

Homogenous catalysis

Reactant(s) and catalyst are in the same phase

Activation energy

A catalyst works by lowering this.

Adsorbed

What happens to a gaseous reactant when it interacts with a solid catalyst

Intermediate compound formation

Mechanism for the reaction of potassium sodium tartarate with hydrogen peroxide using cobalt chloride as catalyst